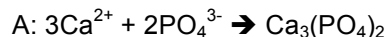
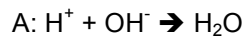


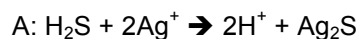
Net Ionic Equation Writing in 2014 and beyond – answers to examples



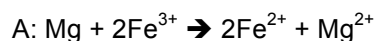
A: Sodium and chloride ions



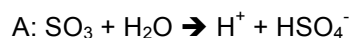
A: $\text{LiBr}_{(\text{aq})}$, neutralization



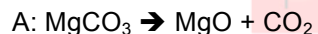
A: Less than 7 since one of the products is H^+ , an acid



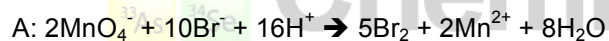
A: Yellow/Brown \rightarrow Pale green as $\text{Fe}^{3+}_{(\text{aq})}$ is converted to $\text{Fe}^{2+}_{(\text{aq})}$



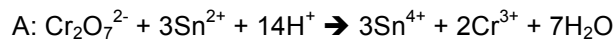
A: Sulfur dioxide



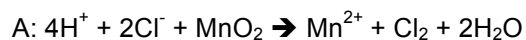
A: Carbon dioxide will turn a solution of limewater "milky"



A: 7+ before \rightarrow 2+ after



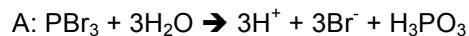
A: Tin(II) ions are losing electrons and become tin(IV)



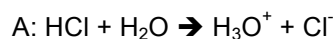
A: Chloride ions. They lose electrons and are oxidized in the process



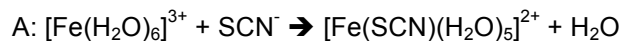
A: Hydrogen gas will give a "squeaky pop" with a lighted splint



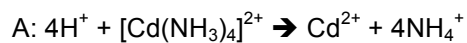
A: Trigonal pyramidal (3BP, 1LP)



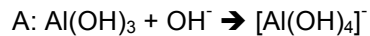
A: HCl and Cl^- or H_3O^+ and H_2O



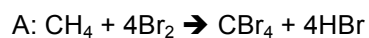
A: Dative bond



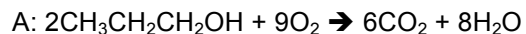
A: N has 4 bonding pairs around it in NH_4^+ so it is sp^3 hybridized



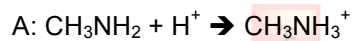
A: It would cause the reaction to shift to the reactants side and to precipitate more solid



A: A substitution. H atoms are replaced by Br atoms



A: 12



A: Acid/Base



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