

AP LAB 00b: Inorganic Nomenclature

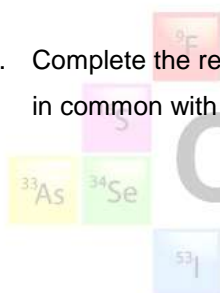
Aim To identify inorganic compounds by name and formula

Apparatus 24 sealed test tubes containing various inorganic compounds. Label with formula or name for each tube.

Chemicals Various

Method

1. Visit each set of chemicals and read the name or formula given on the card for each bottle in the set.
2. Record the name or formula in your results table and fill in the missing name or formula for each bottle in the set.
3. Write a brief description of each compound in the set.
4. Complete the results table by finding something that all three compounds in each set have in common with one another.



Chemistry Pages

Results

SET A		
NAME	FORMULA	DESCRIPTION
1.		
2.		
3.		
Commonality		

SET B		
NAME	FORMULA	DESCRIPTION
1.		
2.		
3.		
Commonality		

SET C		
NAME	FORMULA	DESCRIPTION
1.		
2.		
3.		
Commonality		

SET D		
NAME	FORMULA	DESCRIPTION
1.		
2.		
3.		
Commonality		

SET E		
NAME	FORMULA	DESCRIPTION
1.		
2.		
3.		
Commonality		

SET F		
NAME	FORMULA	DESCRIPTION
1.		
2.		
3.		
Commonality		

SET G		
NAME	FORMULA	DESCRIPTION
1.		
2.		
3.		
Commonality		

SET H		
NAME	FORMULA	DESCRIPTION
1.		
2.		
3.		
Commonality		

Conclusion/Calculation


1. Describe a feature that all ionic compounds have in common.


2. Ionic compounds are usually formed between which types of elements?


3. Complete the table below.


GROUP	1	2	13	15	16	17
CHARGES ON IONS						

4. What type of elements exhibit the tendency to form metal ions of more than one positive charge? How are these different charges denoted when naming compounds?

<p>A1</p> <p>Sodium nitrate</p>	<p>A2</p> <p>K_2SO_4</p>	<p>A3</p> <p>Lithium chloride</p>
<p>B1</p> <p>$CuCl_2$</p> 	<p>B2</p> <p>$CuSO_4$</p>	<p>B3</p> <p>Copper(II) nitrate</p>

<p>C1</p> <p>Copper(II) sulfate</p>	<p>C2</p> <p>Cobalt(II) chloride</p>	<p>C3</p> <p>$\text{Ca}(\text{NO}_3)_2$</p>
<p>D1</p> <p>Copper(II) chloride</p> 	<p>D2</p> <p>Zinc oxide</p>	<p>D3</p> <p>Chromium(III) nitrate</p>

<p>E1</p> <p>KClO_3</p>	<p>E2</p> <p>KMnO_4</p>	<p>E3</p> <p>$\text{Co}(\text{NO}_3)_3$</p>
<p>F1</p> <p>FeCl_3</p> 	<p>F2</p> <p>Sodium fluoride</p>	<p>F3</p> <p>NaCl</p>

<p>G1</p> <p>Hydrochloric acid</p>	<p>G2</p> <p>H_2SO_4</p>	<p>G3</p> <p>NH_4OH</p>
<p>H1</p> <p>$\text{Ca}(\text{OH})_2$</p> 	<p>H2</p> <p>Sodium hydroxide</p>	<p>H3</p> <p>Sodium carbonate</p>